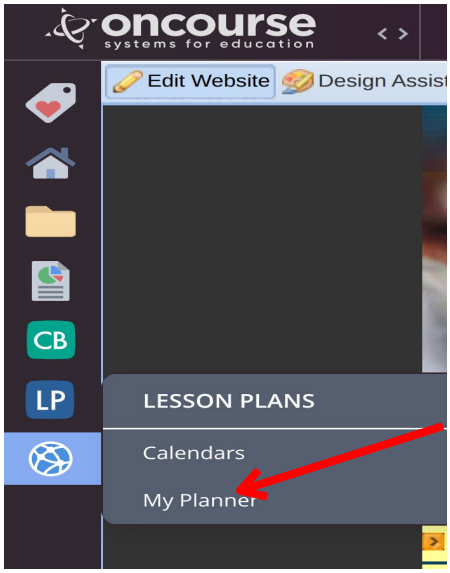
2026 OnCourse Lesson Plans

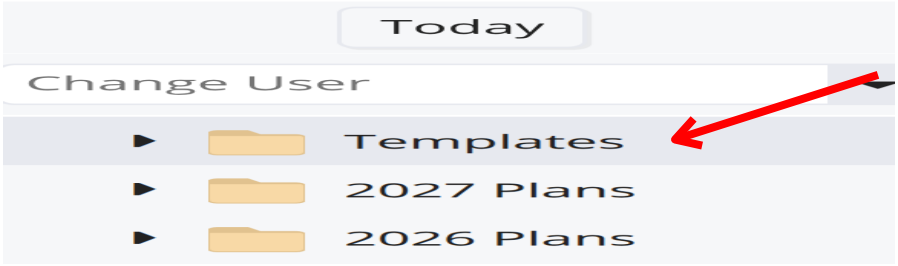
– Vanessa Small & Jodie Moorhead

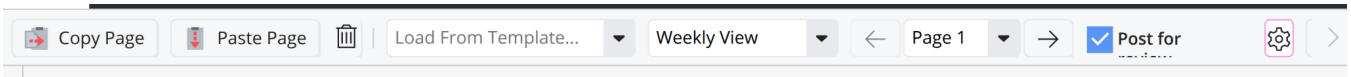


I. **Log in to OnCourse (https://www.oncoursesystems.com)**
   - **a.** Username – same as your email
   - **b.** Password – see Jodie

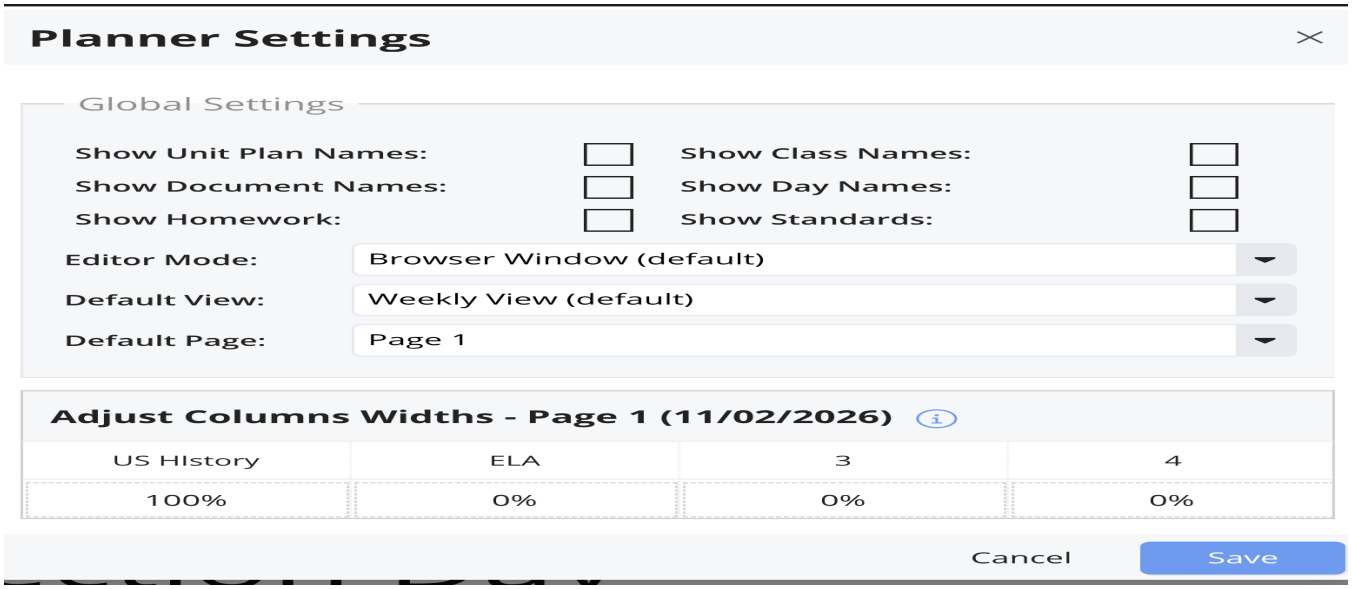
II. **Getting to Your Lesson Plans**
   - **a.** On the homepage of OnCourse, click the tab at the top "Lesson Plans".
   - **b.** Then click "My Planner".
   - **c.**

III. **Creating Your Template**
   - **a.** Before you make your lesson plans, it may be easier to create a "Template" first.
   - **b.** On the left–hand side of the page, click the Template folder icon and then click "Add Template".



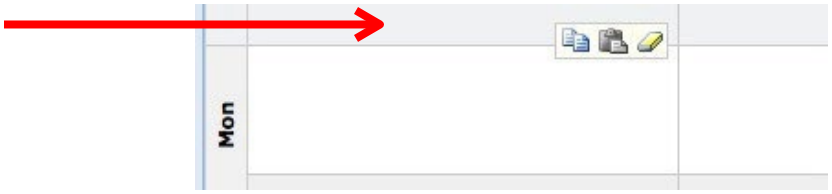


   - **c.** You can break the Template into 1, 2, 3, or 4 sections. To do this, click the "Settings" button.



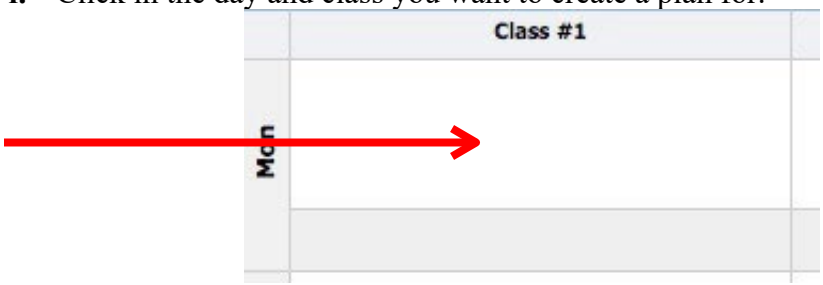
d. Once in the settings, change the percentages to what you want (make sure the values add up to 100%)

- For example, if you want two columns, change the percentages in 1 and 2 to 50% and change 3 and 4 to 0%.
- Or if you want three columns, change 1, 2, and 3 to 33% and change 4 to 0%.
- Next to name your classes, click in the blank area located on the top of one of the columns.



e. After you name each class you want lesson plans for, you can add whatever you want to your lesson plan.

f. Click in the day and class you want to create a plan for.



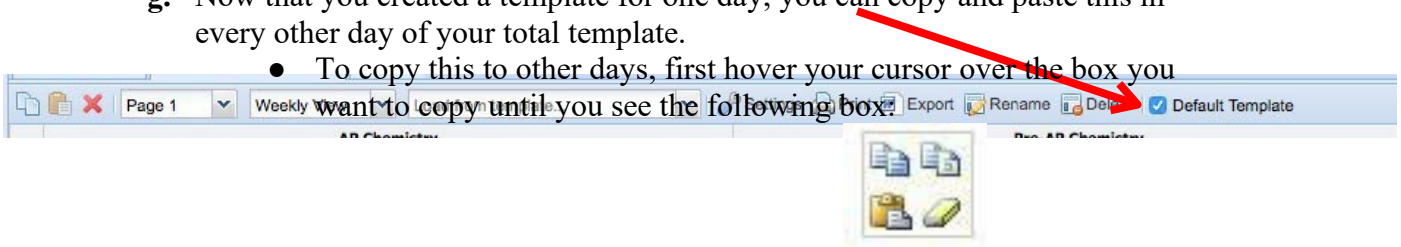
- Here you can edit your plans just as you would a Word Document.
- These lesson plans do not have to be extremely detailed.
- They need to show you and the administration what you will be doing in class that day.
- What you put in here is up to you.
  - ➢ Here is a copy of what I have in all of my lesson plan boxes (this is my template):



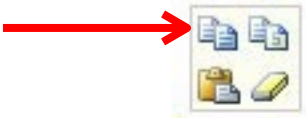
- I include this template for all of my classes and just edit what I may be doing that day/unit for each class.

g. Now that you created a template for one day, you can copy and paste this in every other day of your total template.

- To copy this to other days, first hover your cursor over the box you want to copy until you see the following box:



- Then you will click the button to copy a box.



- Once you have the box copied, you can paste it into all the other days (click the paste button).



- To make this your template for all your future lesson plans, check the box at the top that states "Default Template". This way, whenever you make lesson plans for a week, you'll have the option to use this template.

IV. Making Your Lesson Plans

a. Now that you have a template, you can make your lesson plans.

b. Click the calendar to the left to make lesson plans for your desired week.

- It will ask if you want to use your "Template" for the week – click Yes.

c. Click on a day to enter the information you want for that day.

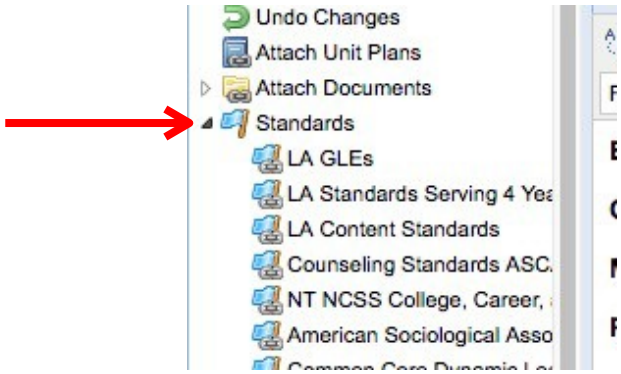
- I use the same Objectives, Materials, and Procedure for an entire unit, but you can change it day to day if you prefer.

d. You can also now add Homework Assignments to corresponding days.

- Click in the "Enter Homework Here" section of the Lesson Plan.
- Enter the Homework that is assigned that day.

e. You can also add Standards to your Lesson Plan.

- Click the "Standards" flag on the left side of the window of the day's lesson plan.
- Choose the Standards you need for each day/unit.



f. After you are done with a day, you need to click the "Save" button to save the information you entered.

g. You can copy the Standards you entered in one day to another day by clicking on the "Copy the Standards In This Block" button and pasting them into the day you want.

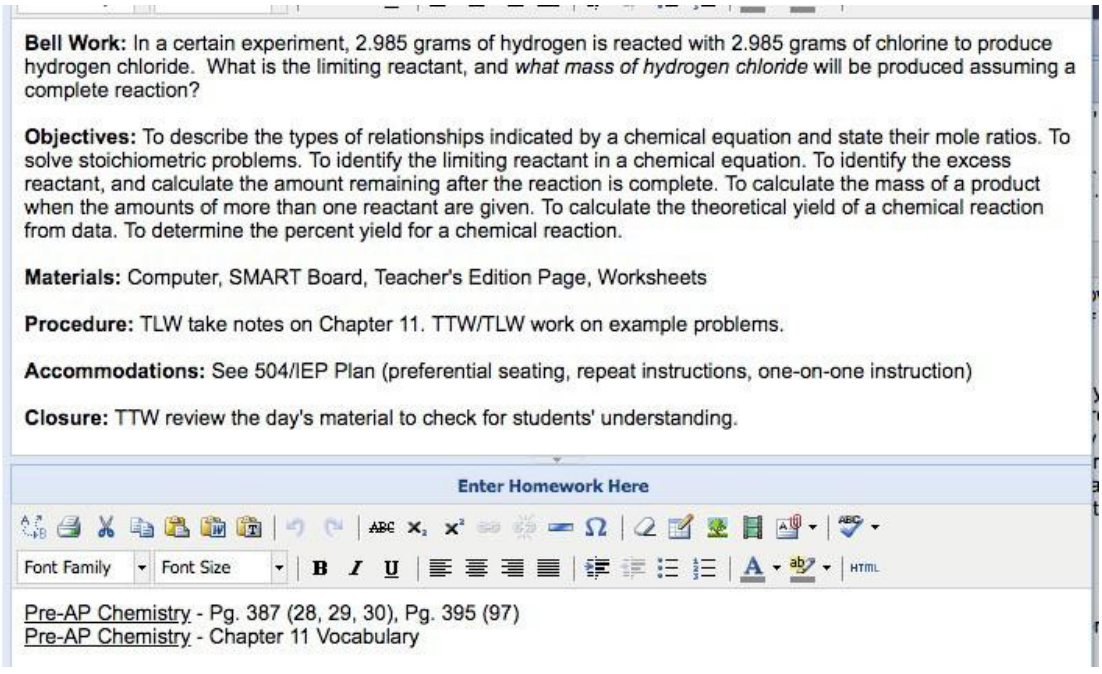


h. You can also now copy and paste Lesson Plan days to other days if you are teaching the same thing multiple days (use the copy and paste button as above).

i. Below are examples of one day and one week of my lesson plans to give you an idea of what it can look like.

V. Resources

a. If you have any questions, contact me (karari.hanks@mcschools.net, x1279) or Jodie (jodie.moorhead@mcschools.net), or you can call OnCourse Technical Support at 1.800.899.7204.



Bell Work: In a certain experiment, 2.985 grams of hydrogen is reacted with 2.985 grams of chlorine to produce hydrogen chloride. What is the limiting reactant, and *what mass of hydrogen chloride* will be produced assuming a complete reaction?

Objectives: To describe the types of relationships indicated by a chemical equation and state their mole ratios. To solve stoichiometric problems. To identify the limiting reactant in a chemical equation. To identify the excess reactant, and calculate the amount remaining after the reaction is complete. To calculate the mass of a product when the amounts of more than one reactant are given. To calculate the theoretical yield of a chemical reaction from data. To determine the percent yield for a chemical reaction.

Materials: Computer, SMART Board, Teacher's Edition Page, Worksheets

Procedure: TLW take notes on Chapter 11. TTW/TLW work on example problems.

Accommodations: See 504/IEP Plan (preferential seating, repeat instructions, one-on-one instruction)

Closure: TTW review the day's material to check for students' understanding.

	AP Chemistry	Pre-AP Chemistry
Mon 09/24/18	**Objectives:** Calculate the wavelength of electromagnetic radiation given its frequency, or its frequency given its wavelength. To order the common kinds of radiation in the electromagnetic spectrum. Understand the concept of photons and calculate their energies. Understand how the uncertainty principle limits how precisely we can specify the position and momentum of electrons. To understand how quantum numbers relate to the number and type of orbitals. To draw energy-level diagrams for the orbitals. Use the periodic table to write electron configurations. **Materials:** Computer, SMART Board, Teacher's Edition Page, Worksheets **Procedure:** TTW/TLW discuss material from Chapter 6. TTW/TLW work on example problems from the chapter. **Accommodations:** See 504/IEP Plan (preferential seating, repeat instructions, one-on-one instruction) **Closure:** TTW review the day's material to check for students' understanding. Homework: AP Chemistry - Pgs. 250-251 (63, 65, 67, 69, 71, 73)	**Bell Work:** Given the data below, calculate the atomic mass of an unknown Element X (***round your answer to the nearest 0.001 amu***). Then, identify the unknown element. **Objectives:** To review material from Chapter 4. **Materials:** Computer, SMART Board, Teacher's Edition Page, CPS Clickers **Procedure:** TTW/TLW review of material from Chapter 4. **Accommodations:** See 504/IEP Plan (preferential seating, repeat instructions, one-on-one instruction) **Closure:** TTW review the day's material to check for students' understanding. Homework: Pre-AP Chemistry - Study for Chapter 4 Test
Tue 09/25/18	**Objectives:** Calculate the wavelength of electromagnetic radiation given its frequency, or its frequency given its wavelength. To order the common kinds of radiation in the electromagnetic spectrum. Understand the concept of photons and calculate their energies. Understand how the uncertainty principle limits how precisely we can specify the position and momentum of electrons. To know how quantum numbers relate to the number and type of orbitals. To draw energy-level diagrams for the orbitals. Use the periodic table to write electron configurations. **Materials:** Computer, SMART Board, Teacher's Edition Page, Worksheets **Procedure:** TTW/TLW discuss material from Chapter 6. TTW/TLW work on example problems from the chapter. **Accommodations:** See 504/IEP Plan	**Bell Work:** Did you study & how much time? **Objectives:** To test on Chapter 4 material. **Materials:** Computer, SMART Board, Tests. **Procedure:** TLW take a test on Chapter 4. **Accommodations:** See 504/IEP Plan (preferential seating, repeat instructions)

	(preferential seating, repeat instructions, one-on-one instruction) **Closure:** TTW review the day's material to check for students' understanding. Homework: AP Chemistry - Pgs. 247-251 (12, 16, 22, 24, 30, 36, 46, 50, 52, 54, 62, 64, 66, 68, 70, 72, 74)	
Wed 09/26/18	**Objectives:** Understand the meaning of effective nuclear charge and how it depends on nuclear charge and electron configuration. To use the periodic table to predict trends in atomic radii, ionic radii, ionization energy, and electron affinity. Write the electron configuration of ions. To understand how atomic properties, such as ionization energy and electron configuration, relate to chemical reactivity and physical properties. **Materials:** Computer, SMART Board, Teacher's Edition Page, Worksheets **Procedure:** TTW/TLW discuss material from Chapter 7. TTW/TLW work on example problems from the chapter. **Accommodations:** See 504/IEP Plan (preferential seating, repeat instructions, one-on-one instruction) **Closure:** TTW review the day's material to check for students' understanding.	**Bell Work:** How many electrons does an atom of neon have? **Objectives:** To compare the wave & particle natures of light. Explain how a quantum of energy relates to an energy change in matter. To compare the Bohr & quantum mechanical models of the atom. To explain de Broglie's wave-particle duality & the Heisenberg uncertainty principle. To identify an atom's energy levels, sublevels, and atomic orbitals. To apply the Pauli exclusion principle, the Aufbau principle, and Hund's rule to write electron configurations and orbital diagrams. **Materials:** Computer, SMART Board, Teacher's Edition Page, Worksheets **Procedure:** TLW take notes on Chapter 5. **Accommodations:** See 504/IEP Plan (preferential seating, repeat instructions, one-on-one instruction) **Closure:** TTW review the day's material to check for students' understanding.
Thu 09/27/18	**Objectives:** Understand the meaning of effective nuclear charge and how it depends on nuclear charge and electron configuration. Use the periodic table to predict trends in atomic radii, ionic radii, ionization energy, and electron affinity. Write the electron configuration of ions. To understand how atomic properties, such as ionization energy and electron configuration, relate to chemical reactivity and physical properties. **Materials:** Computer, SMART Board, Teacher's Edition Page, Worksheets **Procedure:** TTW/TLW discuss material from	**Bell Work:** Label the parts of the wave shown in the picture below. **Objectives:** To compare the wave & particle natures of light. Explain how a quantum of energy relates to an energy change in matter. To compare the Bohr & quantum mechanical models of the atom. To explain de Broglie's wave-particle duality & the Heisenberg uncertainty principle. To identify an atom's energy levels, sublevels, and atomic orbitals. To apply the Pauli exclusion principle, the Aufbau principle, and Hund's rule to write electron configurations and orbital diagrams.

	Chapter 7. TTW/TLW work on example problems from the chapter. **Accommodations:** See 504/IEP Plan (preferential seating, repeat instructions, one-on-one instruction) **Closure:** TTW review the day's material to check for students' understanding. Homework: AP Chemistry - Pgs. 289-290 (11, 15, 19, 23, 25, 29, 31, 35)	**Materials:** Computer, SMART Board, Teacher's Edition Page, Worksheets **Procedure:** TLW take notes on Chapter 5. **Accommodations:** See 504/IEP Plan (preferential seating, repeat instructions, one-on-one instruction) **Closure:** TTW review the day's material to check for students' understanding. Homework: Pre-AP Chemistry - Pg.140 (1, 2, 3); Pg.143 (5, 7); Pg.166 (52, 53); Pg.978 (1, 2, 3, 4)
Fri 09/28/18	**Objectives:** Understand the meaning of effective nuclear charge and how it depends on nuclear charge and electron configuration. To use the periodic table to predict trends in atomic radii, ionic radii, ionization energy, and electron affinity. Write the electron configuration of ions. To understand how atomic properties, such as ionization energy and electron configuration, relate to chemical reactivity and physical properties. **Materials:** Computer, SMART Board, Teacher's Edition Page, Worksheets **Procedure:** TTW/TLW discuss material from Chapter 7. TTW/TLW work on example problems from the chapter. **Accommodations:** See 504/IEP Plan (preferential seating, repeat instructions, one-on-one instruction) **Closure:** TTW review the day's material to check for students' understanding.	**Bell Work:** What happens as electrons fall back down to a more stable state? **Objectives:** To compare the wave & particle natures of light. Explain how a quantum of energy relates to an energy change in matter. To compare the Bohr & quantum mechanical models of the atom. To explain de Broglie's wave-particle duality & the Heisenberg uncertainty principle. To identify an atom's energy levels, sublevels, and atomic orbitals. To apply the Pauli exclusion principle, the Aufbau principle, and Hund's rule to write electron configurations and orbital diagrams. **Materials:** Computer, SMART Board, Teacher's Edition Page, Worksheets **Procedure:** TLW take notes on Chapter 5. **Accommodations:** See 504/IEP Plan (preferential seating, repeat instructions, one-on-one instruction) **Closure:** TTW review the day's material to check for students' understanding.